Centre Number

2

wjec

First name(s)

**GCE A LEVEL** 

1400U50-1E

FRIDAY, 28 APRIL 2023 – MORNING

### BIOLOGY – A2 unit 5 Practical Examination Practical Analysis Task

1 hour

For Examiner's use only					
Question	Maximum Mark	Mark Awarded			
1.	20				
2.	10				
Total	30				

### ADDITIONAL MATERIALS

In addition to this examination paper, you will need a calculator and a ruler.

### INSTRUCTIONS TO CANDIDATES

Use black ink or black ball-point pen. Do not use gel pen or correction fluid.

Pencil may be used to draw tables and graphs.

Write your name, centre number and candidate number in the spaces at the top of this page. Write your answers in the spaces provided in this booklet.

### INFORMATION FOR CANDIDATES

The number of marks is given in brackets at the end of each question or part-question. The maximum mark for this paper is 30.

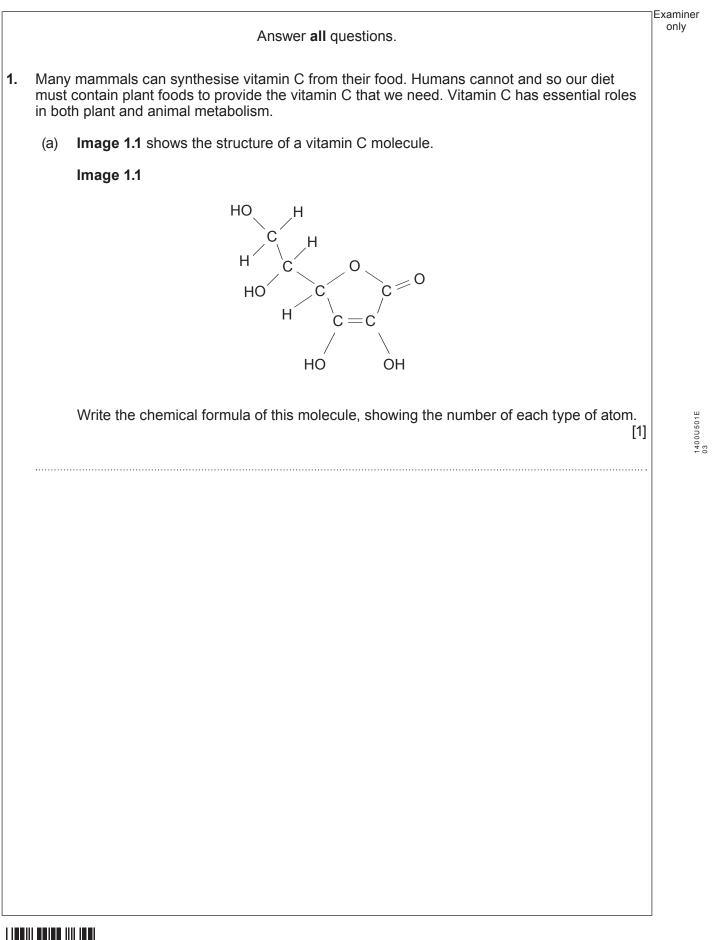


# **BLANK PAGE**

2

## PLEASE DO NOT WRITE ON THIS PAGE







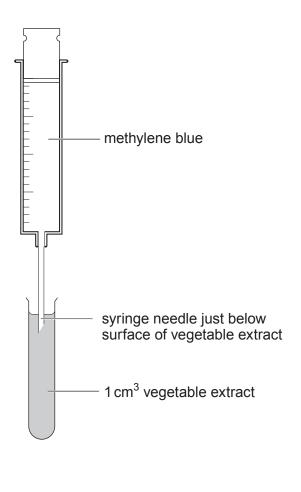
(b) Methylene blue is blue when oxidised and colourless when reduced. Vitamin C is a reducing agent which can reduce methylene blue and decolourise it. The more vitamin C that is present, the greater the volume of methylene blue that can be decolourised. If you know the concentration and volume of methylene blue that is decolourised by a vegetable extract solution, you can find the concentration of vitamin C that the extract contains.

Examiner only

Steps 1–3 and Image 1.2 outline the technique:

- 1. Place  $1 \text{ cm}^3$  of a vegetable extract containing vitamin C in a small container.
- 2. Fill a syringe fitted with a needle with 0.1 mg 100 cm<sup>-3</sup> methylene blue. Slowly inject the methylene blue just below the surface of the vegetable extract. Take great care not to disturb the liquid surface, to avoid introducing air into the solution.
- 3. Continue to add methylene blue until it is no longer reduced i.e. it remains blue when injected. Read the volume of methylene blue that has been injected.

Image 1.2





(i) **Box 1.3** shows an extract from the entry on a Student Safety Sheet for using methylene blue:

5

### Box 1.3

WARNING may cause allergic reaction on contact with skin; causes skin/serious eye irritation;

Use the information given in **Box 1.3** to complete **Table 1.4**, which is a risk assessment for methylene blue.

#### Table 1.4

Ha	azard	Risk	Control measure
(ii)	Identify one oth	er hazard in this experiment <b>an</b>	d state the risk associated wi
•••••			
•••••			
(iii)	If the liquid surf	ace is disturbed while the meth	wlene blue is being injected
(111)	enters the vege	table extract being tested. Sug	gest what effect this might h
	the volume of m	nethylene blue that would need	to be added. Explain your a
•••••			
·····			



Turn over.

Examiner only

[1]

(c) **Table 1.5** shows the results of 15 trials of an experiment to measure the vitamin C content of peas and of green cabbage.

Vitamin C content/mg 100 g <sup>-1</sup>				
Peas	Green cabbage			
35	40			
36	35			
43	34			
47	29			
40	36			
32	36			
38	37			
46	28			
46	41			
41	42			
39	34			
31	34			
45	36			
44	38			
37	37			
Mean = 40.0	Mean =			

### Table 1.5

Calculate the mean concentration of vitamin C in green cabbage. Write your answer in Table 1.5. [1]

Space for working



Examiner only

(d)		atistical table showed that the critical value of t at $p = 0.05$ and 28 degrees of dom is 2.048.	
	(i)	Explain the meaning of ' $p = 0.05$ ' in this experiment.	[1]
	(ii)	Show how, in this analysis, the number of degrees of freedom is calculated to be 28.	[1]
(e)	(i)	The Student's t-test was used to compare the concentrations of vitamin C in per- and green cabbage. A student constructed a null hypothesis for this test, statin 'There is no difference between the vitamin C concentrations of peas and green cabbage.' Rewrite the null hypothesis, showing <b>two</b> ways in which the student's null hypothesis could be improved.	ıg:
	 (ii)	Using the data in <b>Table 1.5</b> , the calculated value of t was 2.527. In relation to the null hypothesis and the critical value, explain why it would be better to eat 100 peas than 100 g of green cabbage.	



(f) In a second experiment, the iron concentration in four samples of peas and green cabbage was measured. **Table 1.6** shows the readings obtained.

### Table 1.6

Iron content/mg 100 g <sup>-1</sup>			
Peas	Green cabbage		
0.45	0.39		
0.46	0.36		
0.52	0.41		
0.49	0.45		
Median =	Median = 0.40		

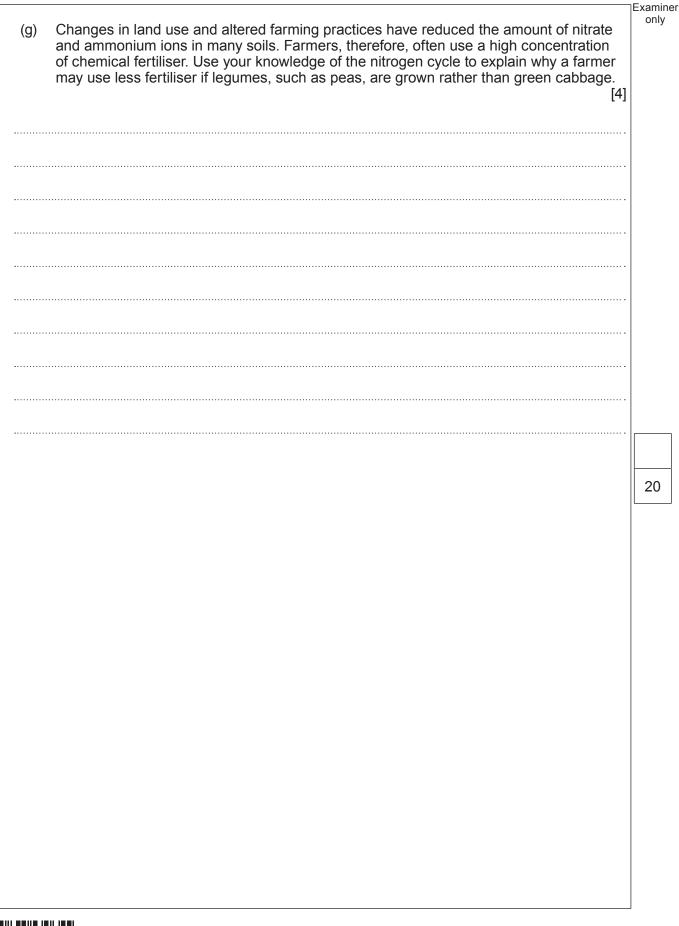
The t-test was not considered to be the most suitable statistical test for comparing these iron concentrations. Instead, a test comparing the median values was used. Find the median value of the concentration of iron in peas. Write your answer in the table.

Space for working

(ii) Describe **one** way this second experiment could be altered so that a t-test could be used to compare the concentrations of iron in peas and green cabbage. [1]

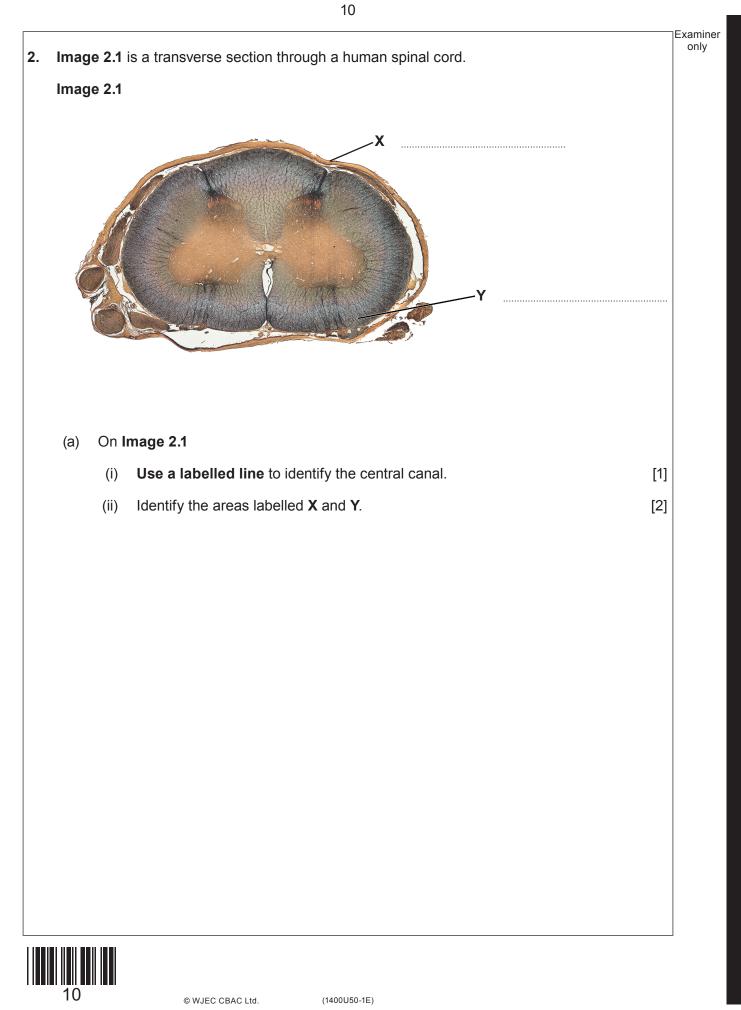


Examiner only

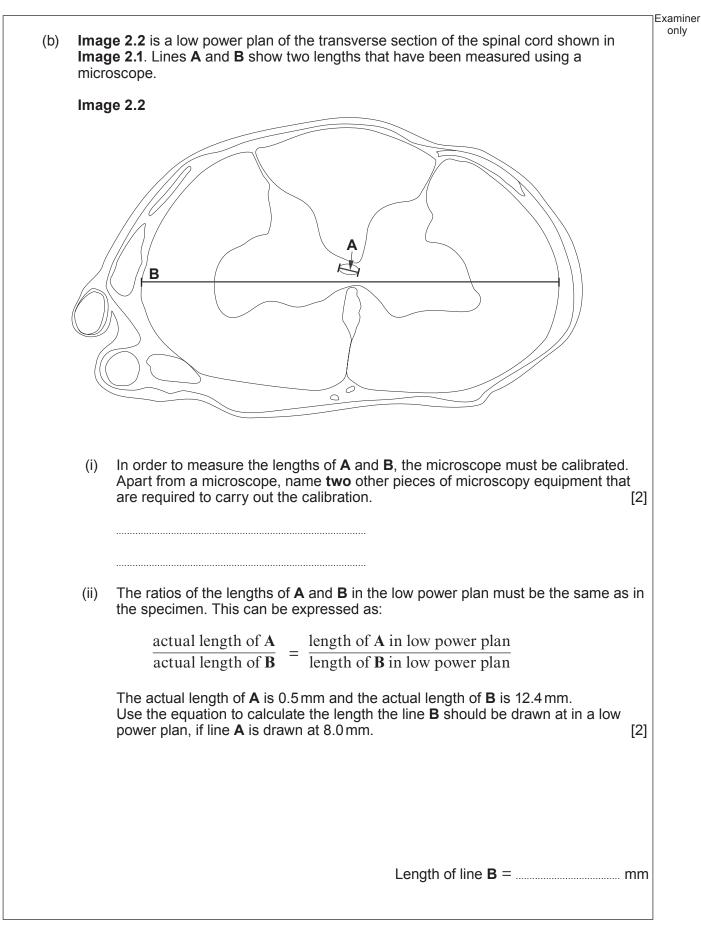




Turn over.



© WJEC CBAC Ltd.





Turn over.

